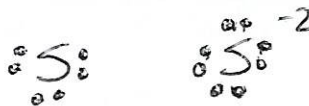
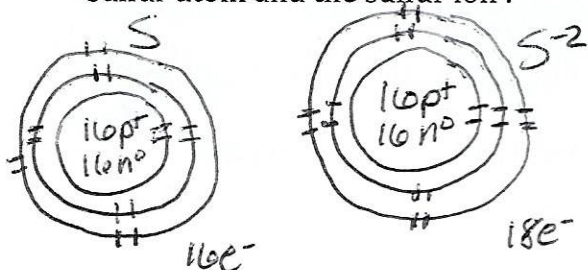


6. Number of valence electrons for Br: 7 Al: 3 Ca: 2 Fr: 1 P: 5 Ar: 8
 7. Most common charge for Br: -1 Al: +3 Ca: +2 Fr: +1 P: -3 Ar: 0
 8. What is the octet rule? atoms gain or lose e⁻ to achieve 8e⁻ in valence shell for stability
 9. What is an ion? charged particle
 10. What is a cation? positively charged particle
 11. What is an anion? negatively charged particle
 12. What is an electron configuration? arrangement of e⁻ around an atom
 13. How many valence electrons does an atom with an electron configuration of $1s^2 2s^2 2p^4$ have? 6
 What is the atom? Oxygen
 14. How many valence electrons does an atom with an electron configuration of $1s^2 2s^2 2p^6 3s^2 3p^2$ have? 4 What is the atom? silicon

Obj: I can draw a Bohr & e-dot model of an element.

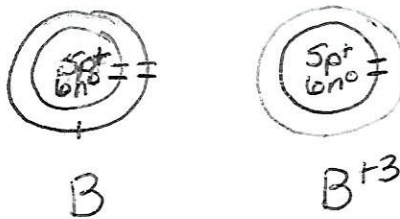
1. a. Draw a Bohr orbital model of both the sulfur atom and the sulfur ion :

- b. Draw Lewis e⁻ Dot diagram of both the sulfur atom and the sulfur ion :



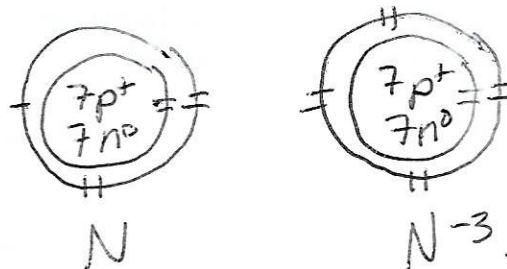
Draw Bohr diagrams of the boron atom and ion. Write the number of each indicated:

2. Total electrons on boron atom 5
 3. Valence electrons on boron atom 3
 4. Common charge on a boron ion +3
 5. Total electrons on a boron ion 2
 6. Valence electrons on a boron ion 2 on 1st energy level



Draw Bohr diagrams of the Nitrogen atom and ion. Write the number of each indicated:

7. Total electrons on nitrogen atom 7
 8. Valence electron on a nitrogen atom 5
 9. Common charge on a nitrogen ion -3
 10. Total electrons on a nitrogen ion 10
 11. Valence electron on nitrogen ion 8



12. How many electrons does a cesium (#55) atom have? 55 cesium ion? 54 Common charge on cesium ion? +1
 13. How many electrons does a aluminum atom have? 13 aluminum ion? 10 Common charge on aluminum ion? +3